STEM Center (BLRC 200)
Academic Success Center
HCC Brandon Campus

Electrochemistry

Oxidation: Loss of electrons, positive charge.

Ex. $Cl \rightarrow Cl^{+2}$ (Cl lost 2 electrons).

Reduction: Gain of electrons, negative charge.

Ex. $O \rightarrow O^{-2}$ (O gained 2 electrons).

Oxidant: The atom being reduced.

Reductant: The atom being oxidized.

$\Delta G = -RT\ln K$

$\Delta G > 0$ non- spontaneous

$\Delta G < 0$ spontaneous

$\Delta G = 0$ Equilibrium

$\Delta G = -nFE_{cell}$

E_{cell}:

- $\Delta G > 0$ spontaneous
- $\Delta G < 0$ non spontaneous